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Chapter 10 "Chemical Quantities" Vocab.  
the SI unit representing  $6.02 \times 10^{23}$   
representative particles of a substance.  
the temperature and pressure at which  
one mole of gas occupies a volume of  
22.4 L. equal volumes of gases at the  
same temperature and pressure contain  
equal numbers of particles.

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## **Quia - Chapter 10 "Chemical Quantities" Vocab**

Chapter 10 "Chemical Quantities". the SI unit representing  $6.02 \times 10^{23}$  representative particles of a substance. the temperature and pressure at which one mole of gas occupies a volume of 22.4 L. equal volumes of gases at the same temperature and pressure contain equal numbers of particles.

## **Quia - Chapter 10 "Chemical Quantities"**

10.2 Mole to Mole & Mole to Volume Relationships \*For all the problems on this page, first find the molar mass of the compound. 1. What is the mass of 9.45 mol of aluminum oxide? mol A ILO + 3 (I A GO q 63,5ù?- 2. What is the mass of  $4.52 \times 10^{-3}$  mol of ethylbenzene,  $C_6H_5CH_2CH_3$ ? \*Ethylbenzene is a hydrocarbon that is produced by burning coal.

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## Answer Key

### CHAPTER 10: Chemical Quantities

**BASICS:**

- The basic unit that is used to determine the amount of a chemical substance is called a mole
- A mole (mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance
- The mole was founded by a scientist named Avagadro, and he decided to use the

### **CHAPTER 10: Chemical Quantities**

10. Look at Figure 10.16 and Table 10.3.

Name three compounds that have an empirical formula of CH. 11. Fill in the labels on the diagram below.

MICROSCOPIC INTERPRETATION

MACROSCOPIC INTERPRETATION

SO<sub>3</sub> molecule 1 mol SO<sub>3</sub> SO<sub>3</sub> composed of composed of S atom and sulfur atoms (  $10^{23}$ ) oxygen atoms 3 atoms and

CHAPTER 10, Chemical  
Quantities (continued)

### **SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)**

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Chapter 10 Chemical Quantities 243  
Section Review Objectives • Convert the mass of a substance to the number of moles of a substance, and the number of moles of a substance to mass • Calculate the volume of a quantity of gas at STP Vocabulary • Avogadro's hypothesis • standard temperature and pressure (STP) • molar volume Key Equations • mass (grams) number of moles

### **Chapter 10 Chemical Quantities Test Answer Key**

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## Answer Key

Section Review Objectives • Convert the mass of a substance to the number of moles of a substance, and the number of moles of a substance to mass • Calculate the volume of a quantity of gas at STP Vocabulary • Avogadro's hypothesis • standard temperature and pressure (STP) • molar volume Key Equations • mass (grams) number of moles

## **Core Teaching Resources Chemistry Answers Chapter 10 ...**

chapter 2/chapter 10/chapter 11  
chemical quantities. day: class activity:  
assignment: worksheets/links/podcasts:  
wed. 11/20. go over test

## **CHAPTER 2/CHAPTER 10/CHAPTER 11**

Chapter 10 "Chemical Quantities" Use this activity to check your knowledge of the vocabulary presented in this chapter

## **Quia - Chapter 10 "Chemical Quantities"**

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## Answer Key

10, that can be made from  $1.09 \times 10^4$  kilograms of phosphorus in the second of these three reactions. The answer is  $2.50 \times 10^4$  kg P<sub>4</sub>O<sub>10</sub>. We used the following steps: Balance chemical equations. (Section 4.1.) Write or identify the definitions of solution, solute, and solvent. (Section 4.2) Given a description of a solution, identify the

### **Chapter 10 Chemical Calculations and equations**

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter Essential Understanding The mole represents a large number of very small particles. Reading Strategy Frayer Model The Frayer Model is a vocabulary development tool. The center of the

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